

Applied Chemistry Notes

Phase Rule

→ **Phase Diagram** :- It is a graph or a plot or graph which gives relationship b/w: various phases in equilibrium as a function of temp, pressure or composition. It is otherwise known as equilibrium diagram.

→ **Type of phase diagram** :- It is of 2 types.

1. Pressure-Temp. diagram (P-T diagram) applicable for one component system. eg:- H_2O system, sulfur system.
2. Temp-composition diagram (T-composition) applicable for two component system. eg:- Bi-Pb system, Fe-C system, Pb-Sn system.

→ **Uses of phase diagram** :-

1. give information about phase stability as a function of temp, pressure and composition.
2. possible to study and to control processes such as solidification of metals and alloys.
3. microstructural changes during heat treatment process.

→ **Phase Rule** :- (Generally known as Gibbs Phase Rule) for any heterogeneous equilibrium system, the equilibrium of the two any no. of phase in a system is not affected or not influenced by gravitational force, electrical force, magnetic force and surface action and only influenced by temperature, pressure and composition.

then $F = C - P + 2$

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Chemistry Wikipedia

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A Brief Review Of Elementary Quantum Chemistry

A Brief Review of Elementary Quantum Chemistry C. David Sherrill School of Chemistry and Biochemistry Georgia Institute of Technology Last Revised on 27 January 2001

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